James Stewart Essential Calculus 2nd Edition

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this book is for instructors who think that most calculus textbooks are too long in writing the book james stewart asked himself what is essential for a three semester calculus course for scientists and engineers essential calculus second edition offers a concise approach to teaching calculus that focuses on major concepts and supports those concepts with precise definitions patient explanations and carefully graded problems the book is only 900 pages two thirds the size of stewart s other calculus texts and yet it contains almost all of the same topics the author achieved this relative brevity primarily by condensing the exposition and by putting some of the features on the book s website stewartcalculus com despite the more compact size the book has a modern flavor covering technology and incorporating material to promote conceptual understanding though not as prominently as in stewart s other books essential calculus features the same attention to detail eye for innovation and meticulous accuracy that have made stewart s textbooks the best selling calculus texts in the world important notice media content referenced within the product description or the product text may not be available in the ebook version

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and mathematics

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this is the the complete student text that includes annotations for teachers as well as wrap around margins in these margins teachers will find notes from co author paul battaglia such as planning guides support resources and more

covering the main fields of mathematics this handbook focuses on the methods used for obtaining solutions of various classes of mathematical equations that underlie the mathematical modeling of numerous phenomena and processes in science and technology the authors describe formulas methods equations and solutions that are frequently used in scientific and engineering applications and present classical as well as newer solution methods for various mathematical equations the book supplies numerous examples graphs figures and diagrams and contains many results in tabular form including finite sums and series and exact solutions of differential integral and functional equations

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this volume introduces techniques and theorems of riemannian geometry and opens the way to advanced topics the text combines the geometric parts of riemannian geometry with analytic aspects of the theory and reviews recent research the updated second edition includes a new coordinate free formula that is easily remembered the koszul formula in disguise an expanded number of coordinate calculations of connection and curvature general fomulas for curvature on lie groups and submersions variational calculus integrated into the text allowing for an early treatment of the sphere theorem using a forgotten proof by berger recent results regarding manifolds with positive curvature

combinatorial enumeration is a readily accessible subject full of easily stated but sometimes tantalizingly difficult problems this book leads the reader in a leisurely way from basic notions of combinatorial enumeration to a variety of topics ranging from algebra to statistical physics the book is organized in three parts basics methods and topics the aim is to introduce readers to a fascinating field and to offer a sophisticated source of information for professional mathematicians desiring to learn more there are 666 exercises and every chapter ends with a highlight section discussing in detail a particularly beautiful or famous result

this book provides an introduction to lie groups lie algebras and repre sentation theory aimed at graduate students in mathematics and physics although there are already several excellent books that cover many of the same topics this book has two distinctive features that i hope will make it a useful addition to the literature first it treats lie groups not just lie alge bras in a way that minimizes the amount of manifold theory needed thus i neither assume a prior course on differentiable manifolds nor provide a con densed such course in the beginning chapters second this book provides a gentle introduction to the machinery of semi simple groups and lie algebras by treating the representation theory of su 2 and su 3 in detail before going to the general case this allows the reader to see roots weights and the weyl group in action in simple cases before confronting the general theory the standard books on lie theory begin immediately with the general case a smooth manifold that is also a group the lie algebra is then defined as the space of left invariant vector fields and the exponential mapping is defined in terms of the flow along such vector fields this approach is undoubtedly the right one in the long run but it is rather abstract for a reader encountering such things for the first time

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Understanding the Ammonia Formula: A Simple Guide

Ammonia, a pungent-smelling gas with a wide range of applications, is a crucial compound in various industries and even plays a vital role in biological processes. Understanding its chemical formula and properties is key to appreciating its significance. This article simplifies the concept of the ammonia formula and explores its related aspects.

1. The Basic Formula: NH₃

The chemical formula for ammonia is simply NH₃. This means each molecule of ammonia consists of one nitrogen atom (N) bonded to three hydrogen atoms (H). This simple formula hides a wealth of chemical and physical properties. The bonds between the nitrogen and hydrogen atoms are covalent bonds, meaning the atoms share electrons to achieve a stable electron configuration. Nitrogen, with five valence electrons, shares three of them with three hydrogen atoms, each contributing one electron to the shared pair. This leaves one lone pair of electrons on the nitrogen atom, which contributes significantly to ammonia's reactivity and properties.

2. Molecular Geometry and Polarity

The NH₃ molecule isn't linear; instead, it adopts a trigonal pyramidal geometry. Imagine a pyramid with the nitrogen atom at the apex and the three hydrogen atoms forming the base. This shape arises because of the lone pair of electrons on the nitrogen atom, which repels the bonding pairs, pushing the hydrogen atoms slightly closer together. This asymmetrical geometry, coupled with the difference in electronegativity between nitrogen and hydrogen, makes ammonia a polar molecule. Nitrogen is more electronegative than hydrogen, meaning it attracts the shared electrons more strongly. This creates a partial negative charge (δ -) on the nitrogen atom and partial positive charges (δ +) on the hydrogen atoms. This polarity is crucial for ammonia's ability to dissolve in water and participate in hydrogen bonding.

3. Ammonia's Properties: A Result of its

Formula

The NH₃ formula directly influences ammonia's properties: Solubility: Ammonia's polarity and ability to form hydrogen bonds with water make it highly soluble in water. This is why ammonia is often found in aqueous solutions. Boiling Point: Relatively low boiling point (-33.34 °C) compared to other compounds of similar molecular weight. This is because hydrogen bonding between ammonia molecules is weaker than in water. Alkalinity: Ammonia acts as a weak base in water, accepting a proton (H+) from water to form ammonium ions (NH₄+) and hydroxide ions (OH-). This is why ammonia solutions are alkaline and have a characteristic pungent smell. Reactivity: The lone pair of electrons on the nitrogen atom makes ammonia a good nucleophile, meaning it can readily donate electrons to other molecules, participating in many chemical reactions.

4. Practical Applications

The versatility of ammonia stems directly from its formula and properties: Fertilizers: Ammonia is a crucial component in the production of nitrogen-based fertilizers like ammonium nitrate and urea, vital for agricultural productivity. The nitrogen atom provides the essential nutrient for plant growth. Cleaning Products: Ammonia's alkaline nature makes it effective in cleaning agents, capable of dissolving grease and grime. Refrigeration: Ammonia's ability to absorb heat during evaporation makes it a useful refrigerant in some industrial applications. Pharmaceuticals: Ammonia is used as a reactant or solvent in the synthesis of various pharmaceuticals. Industrial Processes: It's used in the production of various chemicals, including nitric acid and nylon.

5. Key Insights and Takeaways

Understanding the simple NH₃ formula unlocks a deeper comprehension of ammonia's complex properties and vast applications. Its polarity, geometry, and reactivity are all direct consequences of this seemingly straightforward formula. This knowledge is valuable across various scientific disciplines, from chemistry and biology to agriculture and engineering.

FAQs:

1. Is ammonia dangerous? Yes, anhydrous ammonia is highly toxic and corrosive. Inhaling it can cause severe respiratory problems, and skin contact can lead to burns. Dilute

solutions are less hazardous but should still be handled with care. 2. How is ammonia produced industrially? The Haber-Bosch process is the primary industrial method for ammonia synthesis, reacting nitrogen gas (N_2) and hydrogen gas (H_2) under high pressure and temperature with a catalyst. 3. What is the difference between ammonia and ammonium? Ammonia (NH_3) is a neutral molecule, while ammonium (NH_4^+) is a positively charged ion formed when ammonia accepts a proton. 4. Can ammonia be found naturally? Yes, ammonia is naturally produced by the decomposition of organic matter and is also present in small amounts in the atmosphere. 5. What is the pH of ammonia solution? The pH of an ammonia solution is greater than 7, indicating its alkaline nature. The exact pH depends on the concentration of the ammonia solution.

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